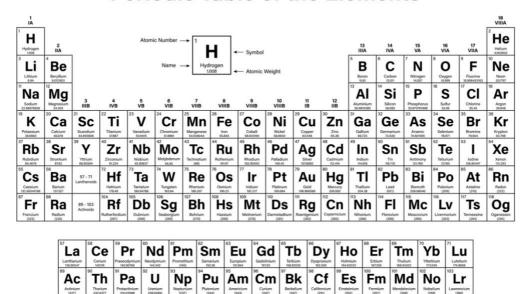
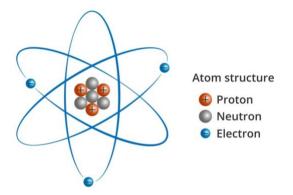


Science V13/S4 - Arrangement of the Periodic Table

Periodic Table of the Elements



You may already be familiar with the periodic table of the elements, but do you know why it looks the way it does? The table is a series of 118 boxes with different symbols and numbers, with each box representing one of the 118 known elements. These boxes are organized into columns and rows. The seven horizontal rows are called periods. The 18 vertical columns are called groups. All of the known chemical elements are listed according to increasing atomic number. The atomic number refers to the number of protons in the nucleus of an atom.



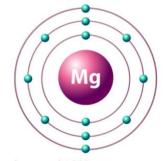
Atoms are made up of protons and neutrons in a nucleus, surrounded by a cloud of electrons that is organized in shells. For example, an atom of Magnesium (Mg) has 12 protons in its nucleus; therefore, the atomic number of Magnesium is 12, and its ranking on the periodic table of chemical elements is 12.

But why aren't all the periods filled in? And how was it decided how many groups there would be? Something very special about the periodic table is that it is



arranged so that elements with similar chemical behaviors are placed together, also called periodic trends. An example of these periodic trends is group 18, the last column of the table, known as the noble gases. All the elements in group 18 are gases at room temperature and pressure. Where elements are located on the table can tell us about how easily they melt, if they are a metal, or how they might bond with other elements.

12 **Magnesium**



Ma

Atomic mass: 24.305 Electron configuration: 2, 8, 2

The reason these trends occur is because of the number of valence electrons. While the protons and neutrons are in the center of the atom in the nucleus, the electrons move about in shells. Each shell can hold a different number of electrons, and once it is full, the electrons start to fill the next shell. You can see this in images of elements that show circles with dots around a nucleus, each circle representing a shell and each dot representing an

electron in that shell. The electrons in the outmost shell are called valence electrons. Valence electrons are important because those are the electrons available to interact with other the other valence electrons of other atoms.

As a general rule, the valence electrons of an element can be determined by the group number in the Periodic Table. Elements with the same number of valence electrons have similar chemical properties. You will notice that, even though the elements in group 2 all have a different number of total electrons, they all have 2 valence electrons, which means that they behave similarly, creating periodic trends.

